

CBSE EXAMINATION PAPER—2019 (Delhi)
CHEMISTRY
(Theory)
Class XII

Time Allowed : 3 hrs.

Maximum Marks : 70

GENERAL INSTRUCTIONS:

- (i) *All questions are compulsory.*
- (ii) *Section A : Questions number 1 to 5 are very short answer questions and carry 1 mark each.*
- (iii) *Section B : Questions number 6 to 12 are short answer questions and carry 2 marks each.*
- (iv) *Section C : Questions number 13 to 24 are also short answer questions and carry 3 marks each.*
- (v) *Section D : Questions number 25 to 27 are long answer questions and carry 5 marks each.*
- (vi) *There is no overall choice. However, an internal choice has been provided in **two** questions of **one** mark, **two** questions of **two** marks, **four** questions of **three** marks and all the **three** questions of **five** marks weightage. You have to attempt only one of the choices in **such** questions.*
- (vii) *Use log tables, if necessary. Use of calculators is **not** allowed.*

Set-I

SECTION A

- 1. Out of NaCl and AgCl, which one shows Frenkel defect and why? (1)
- 2. Arrange the following in increasing order of boiling points: (1)
(CH₃)₃N, C₂H₅OH, C₂H₅NH₂
- 3. Why are medicines more effective in colloidal state? (1)

OR

What is difference between an emulsion and a gel?

- 4. Define ambident nucleophile with an example. (1)
- 5. What is the basic structural difference between glucose and fructose? (1)

OR

Write the products obtained after hydrolysis of lactose.

SECTION B

- 6. Write balanced chemical equations for the following processes: (2)
 - (i) XeF₂ undergoes hydrolysis.
 - (ii) MnO₂ is heated with conc. HCl.

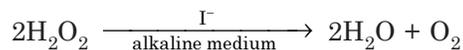
OR

Arrange the following in order of property indicated for each set:

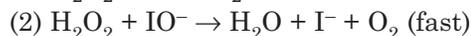
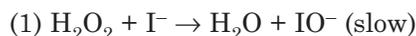
- (i) H₂O, H₂S, H₂Se, H₂Te – increasing acidic character
- (ii) HF, HCl, HBr, HI – decreasing bond enthalpy

(1)

7. State Raoult's law for a solution containing volatile components. Write two characteristics of the solution which obeys Raoult's law at all concentrations.
8. For a reaction (2)



the proposed mechanism is as given below:



(i) Write rate law for the reaction.

(ii) Write the overall order of reaction.

(iii) Out of steps (1) and (2), which one is rate determining step?

9. When MnO_2 is fused with KOH in the presence of KNO_3 as an oxidizing agent, it gives a dark green compound (A). Compound (A) disproportionates in acidic solution to give purple compound (B). An alkaline solution of compound (B) oxidises KI to compound (C) whereas an acidified solution of compound (B) oxidises KI to (D). Identify (A), (B), (C), and (D). (2)
10. Write IUPAC name of the complex $[\text{Pt}(\text{en})_2\text{Cl}_2]$. Draw structures of geometrical isomers for this complex. (2)

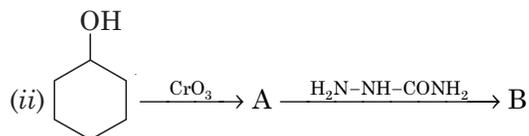
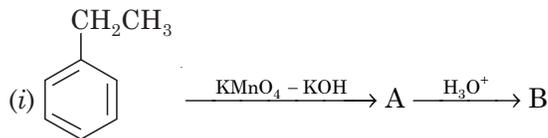
OR

Using IUPAC norms write the formulae for the following:

(i) Hexaamminecobalt(III) sulphate

(ii) Potassium trioxalatochromate(III)

11. Out of $[\text{CoF}_6]^{3-}$ and $[\text{Co}(\text{en})_3]^{3+}$, which one complex is (2)
- (i) paramagnetic
- (ii) more stable
- (iii) inner orbital complex and
- (iv) high spin complex
(Atomic no. of Co = 27)
12. Write structures of compounds A and B in each of the following reactions: (2)

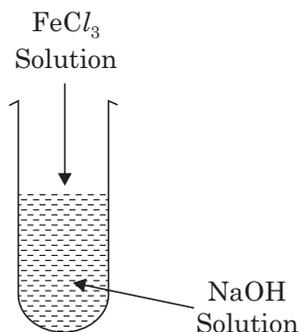


SECTION-C

13. The decomposition of NH_3 on platinum surface is zero order reaction. If rate constant (k) is $4 \times 10^{-3} \text{ Ms}^{-1}$, how long will it take to reduce the initial concentration of NH_3 from 0.1 M to 0.064 M. (3)

(2)

14. (i) What is the role of activated charcoal in gas mask? **(3)**
 (ii) A colloidal sol is prepared by the given method in figure. What is the charge on hydrated ferric oxide colloidal particles formed in the test tube? How is the sol represented?



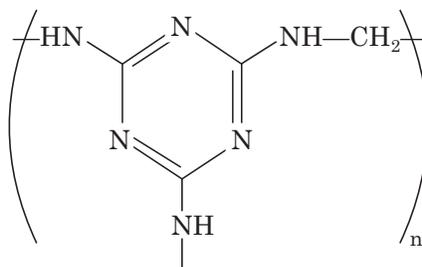
- (iii) How does chemisorption vary with temperature?
15. An element crystallizes in fcc lattice with a cell edge of 300 pm. The density of the element is 10.8 g cm^{-3} . Calculate the number of atoms in 108 g of the element. **(3)**
16. A 4% solution (w/w) of sucrose ($M = 342 \text{ g mol}^{-1}$) in water has a freezing point of 271.15 K. Calculate the freezing point of 5% glucose ($M = 180 \text{ g mol}^{-1}$) in water. (Given: Freezing point of pure water = 273.15 K) **(3)**
17. (a) Name the method of refining which is
 (i) used to obtain semiconductor of high purity,
 (ii) used to obtain low boiling metal.
 (b) Write chemical reactions taking place in the extraction of copper from Cu_2S .
18. Give reasons for the following: **(3)**
 (i) Transition elements and their compounds act as catalysts.
 (ii) E° value for $(\text{Mn}^{2+} | \text{Mn})$ is negative whereas for $(\text{Cu}^{2+} | \text{Cu})$ is positive
 (iii) Actinoids show irregularities in their electronic configuration.
19. Write the structure of monomers used for getting the following polymers: **(3)**
 (i) Nylon-6, 6
 (ii) Glyptal
 (iii) Buna-S

OR

- (i) Is $\left[\text{CH}_2 - \overset{\text{CH}_3}{\text{CH}} \right]_n$ a homopolymer or copolymer? Give reason.

(3)

(ii) Write the monomers of the following polymer:



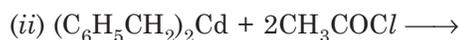
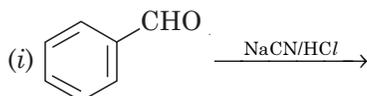
(iii) What is the role of Sulphur in vulcanization of rubber?

20. (i) What type of drug is used in sleeping pills? (3)
 (ii) What type of detergents are used in toothpastes?
 (iii) Why the use of alitame as artificial sweetener is not recommended?

OR

Define the following terms with a suitable example in each:

- (i) Broad-spectrum antibiotics
 (ii) Disinfectants
 (iii) Cationic detergents
21. (i) Out of $(\text{CH}_3)_3\text{C}-\text{Br}$ and $(\text{CH}_3)_3\text{C}-\text{I}$, which one is more reactive towards $\text{S}_{\text{N}}1$ and why? (3)
 (ii) Write the product formed when *p*-nitrochlorobenzene is heated with aqueous NaOH at 443 K followed by acidification.
 (iii) Why *dextro* and *laevo*-rotatory isomers of Butan-2-ol are difficult to separate by fractional distillation?
22. An aromatic compound 'A' on heating with Br_2 and KOH forms a compound 'B' of molecular formula $\text{C}_6\text{H}_7\text{N}$ which on reacting with CHCl_3 and alcoholic KOH produces a foul smelling compound 'C'. Write the structures and IUPAC names of compounds A, B and C. (3)
23. Complete the following reactions: (3)



OR

Write chemical equations for the following reactions:

- (i) Propanone is treated with dilute $\text{Ba}(\text{OH})_2$.
 (ii) Acetophenone is treated with $\text{Zn}(\text{Hg})/\text{Conc. HCl}$.
 (iii) Benzoyl chloride is hydrogenated in presence of Pd/BaSO_4 .

24. Differentiate between the following: (3)

- (i) Amylose and Amylopectin
- (ii) Peptide linkage and Glycosidic linkage
- (iii) Fibrous proteins and Globular proteins

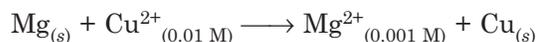
OR

Write chemical reactions to show that open structure of D-glucose contains the following:

- (i) Straight chain
- (ii) Five alcohol groups
- (iii) Aldehyde as carbonyl group

SECTION-D

25. E_{cell}° for the given reaction is 2.71 V (5)



Calculate E_{cell} for the reaction. Write the direction of flow of current when an external opposite potential applied is

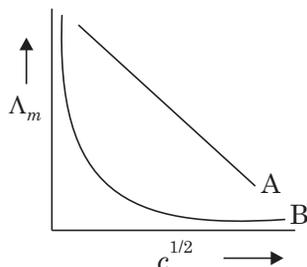
- (i) less than 2.71 V and
- (ii) greater than 2.71 V

OR

(a) A steady current of 2 amperes was passed through two electrolytic cells X and Y connected in series containing electrolytes FeSO_4 and ZnSO_4 until 2.8 g of Fe deposited at the cathode of cell X. How long did the current flow? Calculate the mass of Zn deposited at the cathode of cell Y.

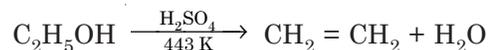
(Molar mass: Fe = 56 g mol⁻¹ Zn = 65.3 g mol⁻¹, 1F = 96500 C mol⁻¹)

(b) In the plot of molar conductivity (Λ_m) vs square root of concentration ($c^{1/2}$), following curves are obtained for two electrolytes A and B:



Answer the following:

- (i) Predict the nature of electrolytes A and B.
 - (ii) What happens on extrapolation of Λ_m to concentration approaching zero for electrolytes A and B?
26. (a) How do you convert the following: (5)
- (i) Phenol to Anisole
 - (ii) Ethanol to Propan-2-ol
- (b) Write mechanism of the following reaction:



(5)

(c) Why phenol undergoes electrophilic substitution more easily than benzene?

OR

(a) Account for the following:

(i) o-nitrophenol is more steam volatile than p-nitrophenol.

(ii) t-butyl chloride on heating with sodium methoxide gives 2-methylpropene instead of t-butylmethylether.

(b) Write the reaction involved in the following:

(i) Reimer-Tiemann reaction

(ii) Friedal-Crafts Alkylation of Phenol

(c) Give simple chemical test to distinguish between Ethanol and Phenol.

27. (a) Give reasons for the following: (5)

(i) Sulphur in vapour state shows paramagnetic behaviour.

(ii) N-N bond is weaker than P-P bond.

(iii) Ozone is thermodynamically less stable than oxygen.

(b) Write the name of gas released when Cu is added to

(i) dilute HNO_3 and

(ii) conc. HNO_3

OR

(a) (i) Write the disproportionation reaction of H_3PO_3 .

(ii) Draw the structure of XeF_4 .

(b) Account for the following:

(i) Although Fluorine has less negative electron gain enthalpy yet F_2 is strong oxidizing agent.

(ii) Acidic character decreases from N_2O_3 to Bi_2O_3 in group 15.

(c) Write a chemical reaction to test sulphur dioxide gas. Write chemical equation involved.

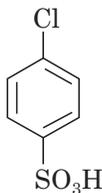
Set-II

SECTION-A

Questions which are different from Set-I.

3. Why conductivity of silicon increases on doping with phosphorus? (1)

5. Write IUPAC name of the given compound: (1)



(6)

SECTION B

8. Write two differences between an ideal solution and a non-ideal solution. (2)

SECTION C

17. (i) Write the role of 'CO' in the purification of nickel. (3)
(ii) What is the role of silica in the extraction of copper?
(iii) What type of metals are generally extracted by electrolytic method?
20. (i) Why bithional is added in soap? (3)
(ii) Why magnesium hydroxide is a better antacid than sodium bicarbonate?
(iii) Why soaps are biodegradable whereas detergents are non-biodegradable?

OR

Define the following terms with a suitable example in each:

- (i) Antibiotics
(ii) Artificial sweeteners
(iii) Analgesics
21. Write the structure of main products when benzene diazonium chloride reacts with the following reagents: (3)
(i) CuCN
(ii) $\text{CH}_3\text{CH}_2\text{OH}$
(iii) KI

Set-III

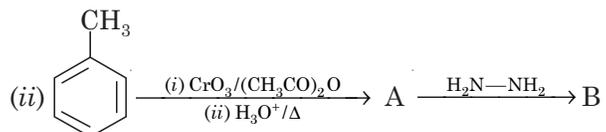
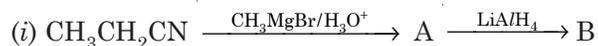
SECTION A

Questions which are different from Set-I and Set-II.

2. What type of stoichiometric defect is shown by ZnS and why? (1)
3. Write one stereochemical difference between $\text{S}_{\text{N}}1$ and $\text{S}_{\text{N}}2$ reactions. (1)

SECTION B

7. State Henry's law and write its two applications. (2)
11. Write the hybridization and magnetic character of following complexes: (2)
(i) $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$
(ii) $[\text{Fe}(\text{CO})_5]$
(Atomic no. of Fe = 26)
12. Write structures of main compounds A and B in each of the following reactions: (2)



SECTION C

17. How will you convert the following: **(3)**
- (i) Impure Nickel to pure Nickel
 - (ii) Zinc blende to Zinc metal
 - (iii) $[\text{Ag}(\text{CN})_2]^-$ to Ag
18. Give reasons for the following: **(3)**
- (i) The transition metals generally form coloured compounds.
 - (ii) E° value for $(\text{Mn}^{3+} | \text{Mn}^{2+})$ is highly positive than that for $(\text{Cr}^{3+} | \text{Cr}^{2+})$ couple.
 - (iii) The chemistry of actinoids elements is not so smooth as that of the lanthanoids.
22. Write equations of the following reactions: **(3)**
- (i) Acetylation of aniline
 - (ii) Coupling reaction
 - (iii) Carbyl amine reaction
24. Define the following with a suitable example in each: **(3)**
- (i) Oligosaccharides
 - (ii) Denaturation of protein
 - (iii) Vitamins

OR

Write the reactions involved when D-glucose is treated with the following reagents:

- (i) Br_2 water
- (ii) $\text{H}_2\text{N}-\text{OH}$
- (iii) $(\text{CH}_3\text{CO})_2\text{O}$